

# Mastering Chemistry Chapter 4 Answers

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### Mastering Chemistry Chapter 4 Answers

4 < 2 < 3 < 1 = 5 The following five beakers, each containing a solution of sodium chloride (NaCl, also known as table salt), were found on a lab shelf: Beaker Contents

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Mastering Chemistry - Chapter 4. For soluble salts, the molarity of a particular ion is equal to the molarity of that compound times the subscript for that ion. For example, 1 M of  $\text{AlCl}_3$  is 1 M in  $\text{Al}^{3+}$  and 3 M in  $\text{Cl}^-$ . 1 M of  $(\text{NH}_4)_2\text{SO}_4$  is 2 M in  $\text{NH}_4^+$  and 1 M in  $\text{SO}_4^{2-}$ .

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Chapter 4 Reading Question 1 Part A What is the elemental symbol for sodium? Hint 1. Elements and chemical symbols All elements are designated by a chemical symbol as found on the periodic table. See Section 4.1 (Page 103) . ANSWER: Correct Chapter 4 Reading Question 2 Part A What element is represented by the symbol K? Hint 1.

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Step 4 of 6 Sodium sulfide : When dissolved in water it forms hydroxide ions therefore, it is alkaline in nature. The presence of sulfide ion can be confirmed by the following test. Step 5 of 6 Sodium chloride : It is ionic in nature so it dissolves in water.

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Chemistry 113 Chapter 4/5 Mastering Chemistry Problem 4.68? Given that one cup = 215.6 mL calculate the molarity of vitamin C in organic juice. Express your answer with the appropriate units. One cup of fresh orange juice contains 113 mg of ascorbic acid (vitamin C,  $C_6H_8O_6$ ).

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